

Chapter 6 Review Chemical Bonding Section 1 Answer Key

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Chapter 6 Review Chemical Bonding

CHAPTER 6 REVIEW Chemical Bonding
SECTION 1 SHORT ANSWER Answer the following questions in the space provided.

- a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures
- b A covalent bond consists of (a) a shared electron.

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hanstep. Key Concepts: Terms in this set
(40) A chemical bond between atoms
results from the attraction between the
valence electrons and _____ of different
atoms. nuclei. A covalent bond consists
of.

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A chemical bond is the force that holds
atoms or ions together as a unit. An
ionic bond is the force that holds cations
and anions together. An ionic bond
forms when electrons are transferred
from one atom to another. Formation of
Ionic Bonds. ... Chapter 6 - Chemical
Bonds

Chapter 6 - Chemical Bonds

CHAPTER 6 REVIEW . Chemical Bonding .

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Answer Key

SHORT ANSWER Answer the following questions in the space provided. 1.

_____ In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms. (d) involved in covalent bonds. 2.

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Review Previous Concepts Chemical Bonding CHAPTER 6 Section 1

Introduction to Chemical Bonding ...

SECTION 6.1 REVIEW 168 CHAPTER 6 (a)

Water molecule, H₂O C₁₂H₂₂O₁₁ (c)

Sucrose molecule, (b) Oxygen molecule,

O₂ Most of the chemicals inside living things and produced by

CHAPTER 6 Chemical Bonding

Modern Chemistry 45 Chemical Bonding

CHAPTER 6 REVIEW Chemical Bonding

SECTION 3 SHORT ANSWER Answer the

following questions in the space

provided. 1. _____ The notation for

sodium chloride, NaCl, stands for one (a)

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formula unit. (c) crystal. (b) molecule.
(d) atom. 2. ____ In a crystal of an ionic
compound, each cation is surrounded by
a ...

CHAPTER 6 REVIEW Chemical Bonding

CHAPTER 6 REVIEW Chemical Bonding
SECTION 6-2 SHORT ANSWER Answer
the following questions in the space
provided. 1. Use the concept of potential
energy to describe how a covalent bond
forms between two atoms. 2. Name two
elements that form compounds that are
exceptions to the octet rule.

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section 1 short answer answer the
following questions in the space
provided. Chapter 6 chemical bonding in
your computer by clicking resolution
image. A a chemical bond between
atoms results from the attraction
between the valence electrons and of
different atoms. B a covalent bond

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consists of a shared electron.

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CHAPTER 6 REVIEW Chemical Bonding
SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the

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Identify and define the three major types of chemical bonding. ... A covalent bond is a chemical bond that involves the sharing of electron pairs between atoms.

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Chapter 6 Test review - Bonding;
Chapter 6 Test Review - Bonding. by
johnsudbury, Nov. 2014. Subjects:
Chemistry ... Chemical Bond: an
attraction between two atoms resulting
from a sharing of outer shell electrons or
the presence of opposite charges on the
atoms.

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CHAPTER 6 REVIEW Chemical Bonding
SECTION 6-3 SHORT ANSWER Answer
the following questions in the space
provided. 1. The notation for sodium
chloride, NaCl, stands for one . (a)
formula unit (c) crystal (b) molecule (d)
atom 2. In a crystal of an ionic
compound, each cation is surrounded by
a number of .

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Chemical Bonding Answer Key for each
of the following types of bonds: a. ionic

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bond b. covalent bond c. metallic bond
2. Use the electronegativity values shown in Figure 5-20, on page 151 of the text, to determine HOMEWORK 6-1
Organic chemistry is study of carbon

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2. b A covalent

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Hybridization delocalizes the bonding electrons over the surface of the entire molecule while molecular orbital theory accounts for bonds one at a time. Hybridization simplifies molecular orbital theory by only considering one bond at a time. Molecular orbital theory for polyatomic molecules considers all of the bonds at once.

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In Chapter 6, we will begin studying how atoms interact with each other to form chemical bonds. Students will review the differences between ionic and covalent bonding and will learn to recognize examples of each, including how to calculate ionic character using electronegativity values.

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